



One mark questions

1. What are two types of matter based on composition?
2. Name two categories of pure substance.
3. Name the types of mixtures.
4. Write the name of any two compounds that sublime on heating.
5. Give two examples of colloids from your daily life.
6. Define the term Solvent.
7. Classify the elements.
8. Name the constituents of German silver.
9. What is the meaning of 'concentration of solution'?
10. What is a condenser?
11. Define dispersion medium.
12. Define Chromatography.
13. Give a natural example of a mixture.
14. Give an example of a liquid and liquid-type solution.
15. What is the principle of separation?
16. What is the meaning of 'Kroma'?
17. Name a metal that is liquid at room temperature.
18. Give the chemical name of chalk and quicklime.
19. Define the term heterogeneous.
20. Name two metals that are highly malleable and ductile.
21. Hydrogen is considered an element. Why?
22. Identify the following as physical or chemical changes:
 - (a) Formation of cloud
 - (b) Water boils to form steam

Three marks questions:

23. Give the difference between mixture and compound.
24. Write the properties of a Solution.
25. Write the properties of a colloidal solution.
26. Draw a flow chart showing the separation of components of Air.
27. What makes water a universal solvent?
28. Differentiate between physical and chemical changes.
29. Melting of wax is a physical change but burning of wax is a chemical change. Explain.
30. Why is water considered a compound?
31. Give some examples where the property: malleability and ductility of metals are used in our life.



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32. Write the characteristics of compounds.
33. How can you test the purity of a given substance?

Five marks questions:

34. What is chromatography? What is its advantage over other methods of separation?
35. Give the difference between true solutions and colloidal solutions.
36. Explain the sublimation process with labelled diagram.
37. Give the difference between colloidal solutions and suspensions.
38. What are colloids? Give their characteristics.
39. Calculate the grams of NaCl (5.25% by mass) in 245 grams of a commercial bleach solution.
40. 110 g of salt is present in 550 g of solution. Calculate the concentration of the solution.
41. How much water should be mixed with 12 ml of alcohol to obtain 12% of alcohol? Calculate.

